



Static Properties of Neutron-Rich and Proton-Rich Isotopes

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Abstract

In this paper, we have investigated some static properties of ^9Be , ^9B , ^{13}C , ^{13}N , ^{17}O , ^{17}F , ^{21}Ne , and ^{21}Na isotopes. The ground state and excited state energy of these isotopes have determined in the relativistic shell model and compared with experimental data. The calculated charge radius of them is in good agreement with experimental results.

Keywords: Charge radius; Energy levels

Introduction

Investigating neutron-rich or proton-rich isotopes are one of the interesting subjects in nuclear science. These isotopes have some single nucleon out of the closed core in shell structure [1]. Determining the energy levels, charge radius and other static properties of nuclei, is one of the useful components to cognition the nuclear structure [2]. Studying the nucleon-nucleon interaction is given many important properties of multi-nucleon systems [3]. The multi-nucleon systems can be investigated in cluster structure [1]. Recently, the cluster model is applied to studying several different systems [3]. These theories have been modeled by mathematic equations like Schrodinger equation in nonrelativistic approach and Dirac equation for spin-1/2 particle and Klein-Gordon equation for spin-1 particles in relativistic approach [4]. Many nucleon-nucleon potentials such as Woods-Saxon potential [5,6], Frost-Musulin potential [7], Pöschl-Teller potential [8] and Eckart potential [1,9] are being used by different researchers.

^9Be is only stable isotopes of beryllium and strongly deformed [10]. The cluster structure of ^9Be is interesting for studying the dynamics of reactions with weakly bound nuclei and also nuclear astrophysics [10]. ^9B has a structure similar to ^9Be , with the odd-neutron replaced by the odd-proton [11].

There are two stable isotopes of carbon: ^{12}C and ^{13}C . ^{13}C is used for instance in organic chemistry research, studies into molecular structures, metabolism, food labeling, air pollution and climate change. It is also used in breath tests to determine the presence of the helicobacter pylori bacteria which causes stomach ulcer. ^{13}C can also be used for the production of the radioisotope ^{13}N which is a PET isotope [12,13].

Natural nitrogen consists of two stable isotopes, ^{14}N , and ^{15}N . Fourteen radioactive isotopes have also been found so far. The ^{13}N decays with a half-life of ten minutes to ^{13}C , emitting a positron [14]. Naturally occurring oxygen is composed of three stable isotopes, ^{16}O , ^{17}O , and ^{18}O with ^{16}O being the most abundant. ^{17}O is a low-abundant, natural, stable isotope of oxygen [15]. Fluorine has 18 known isotopes, with atomic masses ranging from ^{14}F to ^{31}F . ^{19}F is the only stable isotope of fluorine [16].

Sodium is the sixth most common element on Earth [17]. It is used as a heat exchanger in some nuclear reactors, and as a reagent in the chemicals industry. ^{23}Na is the only stable isotopes of sodium [18]. The Earth's crust contains 2.6% sodium, making it the fifth most abundant metal, behind aluminum, iron, calcium, and magnesium [19]. ^{23}Na is created in the carbon-burning process in stars by fusing two carbon atoms together [20].

Neon is a chemical element and the second-lightest noble gas, after helium and very common element in the universe and solar system, but it is rare on earth [17,21]. Stable isotopes of neon are produced in stars. Neon has three stable isotopes, ^{20}Ne , ^{21}Ne and ^{22}Ne [21].

The ^9Be , ^9B , ^{13}C , ^{13}N , ^{17}O , ^{17}F , ^{21}Ne , and ^{21}Na isotopes can be modeled as a doubly magic ($N=Z$) with one additional nucleon out of core [22]. So, these isotopes can be investigated such as single particle in shell model. The experimental energy levels of these isotopes are shown in Figure 1 [23].

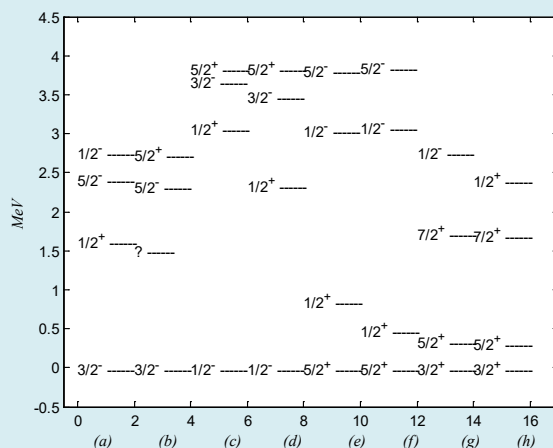


Figure 1: The experimental energy levels of a) ^9Be , b) ^9B , c) ^{13}C , d) ^{13}N , e) ^{17}O , f) ^{17}F , g) ^{21}Ne , h) ^{21}Na .

In this paper, we have calculated the energy levels and charge radius of ^9Be , ^9B , ^{13}C , ^{13}N , ^{17}O , ^{17}F , ^{21}Ne , and ^{21}Na isotopes by solving the Dirac equation using Eckert potential plus coulomb potential for interaction between single nucleon and core cluster.

The Ground State and Excited State Energy of Isotopes

The Dirac equation is one of the most significant equations in physics [24]. There is the exact solution of this equation just only for a few simple interactions. So, the kinds of various methods have been used for the solution of this equation, exemplar, the super symmetric method [24], Nikiforov-Uvarov method [8,25] and so on. By submitting suitable potential in spin symmetry Dirac equation can be written as:

$$\left[\frac{d^2}{dr^2} - \frac{k(k+1)}{r^2} - \frac{1}{\hbar^2 c^2} (Mc^2 + E_{nr,k})(Mc^2 - E_{nr,k} + 8V_0 \frac{e^{-2\alpha r}}{(1 - e^{-2\alpha r})^2} + \frac{2V_1}{r}) \right] F_{nr,k}(r) = 0 \quad (1)$$

In equation (1), V_0 and V_1 are the actual parameter describing the potential well depth and the parameter α representing the potential range.

By suitable approximation and $s = e^{-2\alpha x}$ equation 1 can be

written as equation 2

$$\frac{d^2 F}{ds^2} + \frac{1-s}{s(1-s)} \frac{dF}{ds} + \frac{1}{s^2(1-s)^2} (-\xi_1 s^2 + \xi_2 s - \xi_3) F = 0 \quad (2)$$

Where ξ_i are defined like bellow:

$$\begin{aligned} \xi_1 &= -\frac{E^2 - M^2 C^4}{4\alpha^2 \hbar^2 C^2} \\ \xi_2 &= -2 \frac{E^2 - M^2 C^4}{4\alpha^2 \hbar^2 C^2} - 8 \frac{E + MC^2}{4\alpha^2 \hbar^2 C^2} V_0 + 4 \frac{E + MC^2}{4\alpha^2 \hbar^2 C^2} \alpha V_1 \\ \xi_3 &= k(k+1) - \frac{E^2 - M^2 C^4}{4\alpha^2 \hbar^2 C^2} + 4 \frac{E + MC^2}{4\alpha^2 \hbar^2 C^2} \alpha V_1 \end{aligned} \quad (3)$$

By applying parametric Nikiforov-Uvarov method [22,25-28] the energy Eigen-value formula can be written as

$$\left(n + \frac{l}{2} \right)^2 + \frac{l}{4} + (2n+1) \left(\sqrt{\xi_1 - \xi_2 + \xi_3} + \frac{l}{4} + \sqrt{\xi_3} \right) - \xi_2 + 2\xi_3 + 2 \sqrt{\xi_1 - \xi_2 + \xi_3} + \frac{l}{4} \sqrt{\xi_3} = 0 \quad (4)$$

The ground state and excited state energy of ^9Be , ^9B , ^{13}C , ^{13}N , ^{17}O , ^{17}F , ^{21}Ne , and ^{21}Na isotopes are compared with experimental results in Table 1.

Isotopes	$\alpha(\text{fm}^{-1})$	V_0	V_1	State	$E_{\text{our}}(\text{Mev})$	$E_{\text{exp}}(\text{Mev})$ [29]
9Be	0.0289	0.069	-410.821	1p _{3/2}	-58.1659	-58.164
				1f _{5/2}	-55.7804	-55.735
9B	0.0288	0.07	-414.427	1p _{3/2}	-56.3134	-56.3136
				1f _{5/2}	-53.9537	-53.9686
13C	0.0399	2.2256	-1595.6	1p _{1/2}	-97.1324	-97.1326
				2p _{3/2}	-93.4231	-93.4486
13N	0.0398	2.269	-1610.18	1p _{1/2}	-94.1469	-94.1467
				2p _{3/2}	-90.4895	-90.6447
17O	0.0435	2.8859	-1679.01	1d _{5/2}	-131.7623	-131.7624
				2p _{1/2}	-128.6877	-128.707
17F	0.043	2.902	-1698.53	1d _{5/2}	-128.2352	-128.2353
				2p _{1/2}	-125.155	-125.1313
21Ne	0.0322	0.108	-482.271	1d _{3/2}	-167.404	-167.4047
				1g _{7/2}	-165.2153	-165.6597
21Na	0.032	0.128	-525.442	1d _{3/2}	-163.0454	-163.0463
				1g _{7/2}	-161.1837	-161.3303

Table 1: The ground state and excited state energy of isotopes.

In spin symmetry condition, the upper wave function is achieved in to the form

$$F_{n,k}(r) = N(e^{-2\alpha r})\sqrt{\xi_3} (1 - e^{-2\alpha r}) \sqrt{\xi_1 - \xi_2 + \xi_3 + \frac{1}{4} + \frac{1}{2}} P_n^{(2\sqrt{\xi_3}, 2\sqrt{\xi_1 - \xi_2 + \xi_3 + \frac{1}{4}})} (1 - 2e^{-2\alpha r}) \quad (5)$$

Where N is the normalization constant [30], the lower component of the Dirac spinor can be calculated by equation (6)

$$G_{n,k}(r) = \frac{\hbar^2 c^2}{E + Mc^2} \left(\frac{d}{dr} + \frac{k}{r} \right) F_{n,k}(r) \quad (6)$$

And Wave function for Dirac equation can be calculated from equation (7) as

$$\Psi_{n,k}(r, \theta, \varphi) = \frac{N}{r} \left[\frac{Y_{j,m}^1(\theta, \varphi)}{E + Mc^2} \left(\frac{d}{dr} + \frac{k}{r} \right) Y_{j,m}^1(\theta, \varphi) \right] (e^{-2\alpha r})\sqrt{\xi_3} (1 - e^{-2\alpha r}) \sqrt{\xi_1 - \xi_2 + \xi_3 + \frac{1}{4} + \frac{1}{2}} P_n^{(2\sqrt{\xi_3}, 2\sqrt{\xi_1 - \xi_2 + \xi_3 + \frac{1}{4}})} (1 - 2e^{-2\alpha r}) \quad (7)$$

The charge radius is determined from equation (8)

$$\langle r^2 \rangle^{1/2} = \left(\frac{\int \psi_{n_r,k}^*(r) r^2 \psi_{n_r,k}(r) d^3r}{\int \psi_{n_r,k}^*(r) \psi_{n_r,k}(r) d^3r} \right)^{1/2} \quad (8)$$

The charge radius of isotopes is compared by experimental data in Table 2.

Isotopes	$\langle r^2 \rangle_{0ur}^{1/2} (fm)$	$\langle r^2 \rangle_{exp}^{1/2} (fm)$
⁹ Be	2.5181	2.5190 [30]
⁹ B	2.5378	2.81 [31]
¹³ C	2.4443	2.4614 [30]
¹³ N	2.4718	-----
¹⁷ O	2.6588	2.6932 [30]
¹⁷ F	2.7008	-----
²¹ Ne	2.9665	2.9695
²¹ Na	3.0130	3.0136 [30]

Table 2: The charge radius for ⁹Be, ⁹B, ¹³C, ¹³N, ¹⁷O, ¹⁷F, ²¹Ne, and ²¹Na isotopes for ground state energy.

Conclusion

In this paper, we have considered ⁹Be, ⁹B, ¹³C, ¹³N, ¹⁷O, ¹⁷F, ²¹Ne, and ²¹Na isotopes. Since these isotopes have one additional nucleon out of core, it can be investigated as

single particle model in relativistic shell model. Therefore, we solved the spin symmetry Dirac equation with applying PNU method. By choosing suitable potential for N-cluster interaction, the ground state and excited state energy of isotopes are obtained. These results brought in table1 and compared with experimental data.

The charge radius obtained for ground state ${}^9\text{Be}$, ${}^9\text{B}$, ${}^{13}\text{C}$, ${}^{13}\text{N}$, ${}^{17}\text{O}$, ${}^{17}\text{F}$, ${}^{21}\text{Ne}$, and ${}^{21}\text{Na}$ isotopes. As seen in table 2 our result has good agreement with experimental data.

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